

CHEM B1B | General Chemistry/Chem Analysis

This course was successfully synced with the SIS on 4/27/2026 at 12:19 AM.

COVER INFO

General Information

Cross Listed Courses —

Department(s)

Subject Code - CB01A

CHEM

Department Percent Ownership

Chemistry - 100%

Course Number - CB01B

B1B

eL- Course Title

General Chemistry and Chemical Analysis

Course Short Title

General Chemistry/Chem Analys

eL-Course Description

This course is a continuation of CHEM B1A. The course is intended to provide a broad foundation of chemical principles for science majors. The major topics considered include the properties of solutions and solids, chemical equilibrium, acids and bases, solubility products, thermodynamics, electrochemistry, and nuclear chemistry. The lab portion of the class will include quantitative measurements based on the same topics.

System Taxonomy of Program (TOP) Code - CB03

—

CIP Code

400501 - Chemistry, General. You may refer here to the NCES full list of CIP Codes.

System Taxonomy Code

1905

System Taxonomy Name

Chemistry, General

SAM Priority Code - CB09

—

SAM Priority Name

Non-occupational

SAM Priority Code

E

Minimum Qualifications

Discipline Requiring a Master's Degree (Column 1)

Chemistry

This field requires an entry. Use the Minimum Qualifications Hand to choose all of the Master's degrees listed for this discipline. You can find the hand here: <https://www.cccco.edu> search Minimum Qualifications

Alternate Master's Discipline

—

Do not enter new disciplines in this field. It is here for historical purposes and is being phased out.

Specific Bachelor's/Associate Degree and Professional Experience (Column 2)

—

These disciplines require a specific Bachelor's degree and two years of professional experience, and/or other specifications as noted or specifically named Associate degree(s) and six years of professional experience. Use the Minimum Qualifications Hand to choose degrees listed for this discipline. You can find the hand here: <https://www.cccco.edu> Search Minimum Qualifications

Any Degree and Professional Experience (Column 3)

—

Use the Minimum Qualifications Hand to choose any Bachelor's' degree or higher and two years of professional experience or any associate degree and six years of professional experience listed for this discipline. You can find the hand here: <https://www.cccco.edu> search Minimum Qualifications

Proposal Details

Author(s)

Wade Ellis

Proposal Start

Summer 2025

Effective Start Term

Summer 2025

Notes for Submission

List of updates-Text list was updated to include newer editions of texts.-Precalculus was added as an advisory course.-CSLO #1 was deleted, and the other CSLOs were renumbered.-Course objective 7 was rewritten for clarity.-

Course description rewritten

Attached Files

—

COURSE DEVELOPMENT OPTIONS

Course Options

Course Basic Skill Status - CB08

Course Special Class Status - CB13

— —

Grade Options

Default Grade Mode

Audit (A) Standard Letter (S)

Standard Letter (S)

eL-Course Basic Skills Status eL-Course Special Class Status

Course is not a basic skills course. Course is not a special class.

Retake Policy Description

Non-repeatable credit Course is not a support course

eL-Course Support Course Status

Course Support Course Status - CB26

— —

In-Service

—

Is this course an apprenticeship course?

Competency Based Education (CBE) Course

No

—

Course Prior to College Level - CB21

Course Prior to College Level (CB21)

— Not applicable.

Course Upper Division Status - CB27 [This field is not integrated] [Integrated field as CB21] —

Is this course an approved upper division course associated with a baccalaureate program at the college —

Dependencies

No dependencies

Transferability, C-ID, and General Education Options

Course Transfer Status - CB05

—

eL-Course General Education Status

Y

eL-Transferability

eL-Transferability Status

Transferable to both UC and CSU

Approved

Request for Transferability

No data added

Request for General Education and Transferability

System/Institution

Category

CSU Transfer

CSU Transferable

Status

Approval Date

End Date

Approved

System/Institution

Category

UC Transfer

UC Transferable

Status

Approval Date

End Date

Approved

System/Institution

Category

CSU General Education Breadth

B.1 Physical Universe

Status

Approval Date

End Date

Approved May 1, 2019

System/Institution

Category

CSU General Education Breadth

B.3 Laboratory Activity

Status

Approval Date

End Date

Approved

May 1, 2019

System/Institution

Category

Intersegmental General Education Transfer Curriculum (IGETC)

Area 5.A Physical Sciences

Status

Approval Date

End Date

Approved

System/Institution

Category

Intersegmental General Education Transfer Curriculum (IGETC) Area 5.C Laboratory Science

Status Approval Date End Date
Approved

System/Institution Category

C-ID Chemistry

Status Approval Date End Date
Approved Apr 23, 2015

System/Institution Category

BC General Education Pattern (starting Summer 2025) Area 5: Natural Science

Status Approval Date End Date
Approved

System/Institution Category

California General Education Transfer Curriculum (Cal-GETC) Area 5A: Physical Science

Status Approval Date End Date
Approved

System/Institution Category

California General Education Transfer Curriculum (Cal-GETC) Area 5C: Laboratory

Status Approval Date End Date
Approved

UNITS AND HOURS Credit/Noncredit Options

Course Credit Status - CB04 Course Non-Credit Category - CB22 —
D

eL-Course Credit Status eL-Course Non-Credit Category eL-Non-Credit Characteristics
Credit - Degree Applicable Credit Course. —

Course Classification Status - CB11 — Funding Agency Category - CB23
Y Y

eL-Course Classification Code eL-Funding Agency Category
Credit Course. Not Applicable.

Cooperative Work Experience Education Status - CB10 Variable Credit Course
— No

eL-Cooperative Work Experience Education Status
No

Weekly Student Hours

Auto-calculate fields based on default ratios?

—

eLumen Weekly Student Hours [Non-Integrated]

Lecture Hours In Class Min

3

Laboratory Hours In Class Min

6

Activity Hours In Class Min

0

Lecture Hours Out of Class Min

6

Laboratory Hours Out of Class Min

00

eLumen Weekly Specialty Hours [Non-Integrated]

Description

—

Weekly Student Specialty Hours In Class Min

—

Course Student Hours

eLumen Course Student Hours [Non-Integrated]

Course Duration (Weeks)

18

Total: Lecture Hours In Class Min

54

Total: Lab Hours In Class Min

108

Total: Activity Hours In Class Min

Total: Course In Class (Contact) Hours Min

Total: Lecture Hours Out of Class Min

108

Total: Lab Hours Out of Class Min

0

Total: Activity Hours Out of Class Min

Total: Course Out of Class Hours Min

-

Total Student Learning Hours Min

-

Type

—

Weekly Student Specialty Hours Out of Class Min

—

Hours Per Unit Divisor

54

Total: Lecture Hours In Class Max

54

Total: Lab Hours In Class Max

108

Total: Activity Hours In Class Max

-

Total: Course In Class (Contact) Hours Max

Total: Lecture Hours Out of Class Max

108

Total: Lab Hours Out of Class Max

0

Total: Activity Hours Out of Class Max

0

Total: Course Out of Class Hours Max

-

Total Student Learning Hours Max

-

Credits [Integrated]

Billing Hours

5

Min

Max

—

Operator

—

Min

Max

162

—

Course Units of Credit Minimum - CB07

Course Units of Credit Maximum - CB06

5

—

Operator

—

Min

Max

108

—

Min

Max

54

—

Min

Max

—

—

PRE-REQUISITES AND ENTRANCE SKILLS

Requisites [Populated with eLumen Data]

[ADVANCE REQUIREMENTS](#) [PREREQUISITES](#) [SIMPLE REQUIREMENTS](#)

No Requirement Added

Pre-requisites, Co-requisites, Anti-Requisites, and Advisories

Entrance Skills

No data added

Limitations on Enrollment

Limitations on Enrollment

No data added

SPECIFICATIONS

Methods of Instruction

Methods of Instruction

Method of Instruction

Rationale

[Written Work](#)

—

Method of Instruction

Rationale

[Problem Solving Activity](#)

—

Method of Instruction

Rationale

[Lecture](#)

—

Method of Instruction

Rationale

[Required Reading](#)

—

Method of Instruction

Rationale

[Library Assignment](#)

—

Method of Instruction

Rationale

Group Activities	—
Method of Instruction	Rationale
Laboratory	—
Method of Instruction	Rationale
Field Trip	—
Method of Instruction	Rationale
Computational Work	—
Method of Instruction	Rationale
Discussion	—
Method of Instruction	Rationale
Demonstration	—
Method of Instruction	Rationale
Case Study	—
Method of Instruction	Rationale
Audiovisual Presentations	—

Assignments

Assignments

For the following problems, you are allowed a calculator and a copy of the periodic table. Circle your final answer for each question. Make sure every numerical answer has the correct number of significant figures.

$$\Delta T = K_f m = K_f \times \frac{m}{M} \times 1000 \quad \Delta T = K_b m = K_b \times \frac{m}{M} \times 1000 \quad \Pi = iMRT \quad P = \frac{n}{V}RT \quad P_a = X_a P_{tot} \quad P_b = X_b P_{tot}$$

$$R = 0.08206 \text{ L atm/(mol K)}$$

- What is the boiling point of a solution containing 15.8 g of NaCl dissolved in 350.0 g of water? ($K_b = 0.512 \text{ } ^\circ\text{C kg/mol}$ for water) (3 points)
- Acetone (CH_3COCH_3) has a vapor pressure of 30.8 kPa and diethylether ($\text{CH}_3\text{CH}_2\text{OCH}_2\text{CH}_3$) has a vapor pressure of 71.7 kPa at 25 °C. What is the mole fraction of acetone in a mixture of these two compounds if the mixture has a vapor pressure of 42.0 kPa? (4 points)
- A. 1.0 M NaCl solution is separated from a 1.0 M MgCl₂ solution by a semipermeable membrane. Which of the following accurately describes what will happen to these solutions due to osmosis? (2 points)
 - Water will move from the NaCl solution to the MgCl₂ solution.
 - Water will move from the MgCl₂ solution to the NaCl solution.
 - NaCl will move from the NaCl solution to the MgCl₂ solution.
 - MgCl₂ will move from the MgCl₂ solution to the NaCl solution.
 - There will be no net movement of any kind across membrane.

4. Which statement is NOT TRUE about colligative properties. (1 point)
- A solution always boils at a higher temperature than the pure solvent.
 - A solution always has a lower vapor pressure than the pure solvent.
 - The higher the Van't Hoff factor of a solute, the lower the freezing point of the solution.
 - Osmotic pressure is inversely proportional to temperature.
 - Colligative properties can be used to measure the molar mass of a solute.

Methods of Evaluation

Methods of Evaluation

Assessment Type	Rationale
Homework	—
Assessment Type	Rationale
Written Exams (Quizzes, Midterm, and/or Final Examination)	—
Assessment Type	Rationale
Objective Exams (may include Multiple choice, Matching items, True/false, Essay)	—
Assessment Type	Rationale
Cumulative Final Examination	—
Assessment Type	Rationale
Small Group Lab Assignments	—
Assessment Type	Rationale
Individual Lab Assignments	—
Assessment Type	Rationale
Computational Problem-Solving Demonstrations	—

Equipment

Proper lab equipment is necessary for this course. This includes 12 of the following:

Hot/stir plates, LabQuests (or similar), temperature probes, pH meters, gas pressure sensors, radiation monitors, spectrophotometers, 50 mL burets, 10 mL bulb pipettes, and pipette bulbs.

As well as:

Standard chemistry glassware (i.e., beakers, Erlenmeyer flasks, test tubes, graduated cylinders) are needed for each student.

Other consumables such as disposable pipettes

At least 2 analytical balances, a drying oven, and properly vented fume hoods.

Texts

Texts

Author

Jespersen, N.D., and Hyslop, A..

Title

Publisher

John Wiley & Sons

Date

2022

ISBN

978-1119820642

Author

D. Franz, P. Hooker, G. Kaminski, et al

Title

Chemistry

Publisher

OpenStax

Date

2024

ISBN

978-1-947172-62-3

Author

Burge, Julia

Title

Chemistry, 6th Ed

Publisher

McGraw Hill

Date

2023

ISBN

978-1-264-08577-4

Author

Tro, Nivaldo J.

Title

Chemistry, A Molecular Approach

Publisher

Pearson

Date

2023

ISBN

978-0-13-749361-6

Other Instructional Materials

Other Instruction Materials

Description

Several attached files (all have "lab" in the title) contain samples of labs typically conducted in the Chem B1B laboratory.

Author

—

Citation

—

—Online Resource(s)

Description

Laboratory Experiments for Chemistry: The Central Science (13th Edition) ISBN-13: 978-0321949912

Author

Theodore E. Brown, John H. Nelson, & Kenneth C. Kemp

Citation

—

Online Resource(s)

—

Materials Fee

Materials Fee

—

LEARNING OUTCOMES

Course Objectives

Course Objectives

1. Kinetics: a. Predict and explain the effect of concentration and temperature on reaction rates and determine a rate law from initial rate data.
b. Describe activation energy and calculate its value for a reaction using the Arrhenius equation.
c. Explain the function of homogeneous and heterogeneous catalysis.

- d. Analyze the mechanism of a chemical reaction and how it predicts a rate law., 2. Equilibrium: a. Describe dynamic equilibria in terms of forward and reverse reaction rates.
- b. Use the Law of Mass Action to write equilibrium constant expression for a given reversible reaction and to determine equilibrium constants and concentrations.
- c. Use Le Chatelier's Principle to predict the direction in which equilibrium will shift if a stress is applied., 3. Ionic and Acid-Base Equilibria: a. Write equations representing acid-base and solubility product equilibria.
- b. Calculate equilibrium concentrations of aqueous particles based on acid base and solubility product equilibria.
- c. Characterize weak and strong acids/bases and relevance to pH and other measures.
- d. Relate equilibrium constants and pH to titrations and vice-versa.
- e. Characterize/predict precipitation reactions and apply understanding to separations and qualitative analysis., 4. Thermodynamics: a. Apply the laws of thermodynamics to predict the spontaneity of chemical reactions and their equilibrium constants.
- b. Calculate quantities involving heat, work, internal energy, enthalpy, entropy, and Gibb's free energy changes.
- c. Perform calculation that relates the equilibrium constant to free energy changes in various systems., 5. Electrochemistry and Oxidation-Reduction: a. Describe electrochemical cells both qualitatively and quantitatively.
- b. Perform stoichiometric calculations of electrolytic cells based on Faraday's Law of Electrolysis.
- c. Apply the Nernst equation to determine voltage of non-standard state cells and the concentrations of components., 6. Nuclear Chemistry: a. Calculate binding-energy changes for nuclear processes.
- b. Use the half-life of a nuclide to calculate quantities based on natural decay.
- c. Apply modes of radioactive decay to balance nuclear equations.
- d. Describe nuclear chemistry applications including fission, fusion, transmutation of elements, and the use of radioactive tracers., 7. Organic Chemistry: Utilize correct nomenclature to describe and analyze organic compounds based on their structure and functional group., 8. Descriptive Chemistry: Identify different cations/anions based on solubility, and other physical/chemical properties such as IR spectra and flame emission., 9. Coordination Chemistry: Use crystal-field theory to explain the visible colors of coordination complexes.

CSLOs

Learning Outcome

No data added

Learning Outcomes [Non-Integrated Populated with eLumen Data]

CHEMB1B 1. Upon successful completion of the course, the student will be able to demonstrate proficiency in solving mathematical problems related to chemistry.

CHEMB1B 2. Upon successful completion of the course, the student will be able to apply the concepts of chemical equilibria to explain chemical phenomena (such as acid-base behavior, common ion effect and solubility product).

CHEMB1B 3. Upon successful completion of the course, the student will be able to apply the concepts of thermodynamics to predict chemical phenomena.

CHEMB1B 4. Upon successful completion of the course, the student will be able to apply the concepts of kinetics to chemical reactions.

CHEMB1B 5. Upon successful completion of the course, the student will be able to interpret experimental data according to the scientific method.

ASSESSMENT MAPPING

OUTLINE

Course Outline

Course Outline

Lecture: (54 hours total)

Kinetics (~17.5%)

1. Predict and explain the effect of concentration and temperature on reaction rates.
2. Determine the order of a reaction from initial rate data.
3. Describe activation energy and calculate its value for a reaction using the Arrhenius equation.
4. Explain the function of homogeneous and heterogeneous catalysis

Equilibrium (~10%)

5. Describe dynamic equilibria in terms of forward and reverse reaction rates.
6. Use the Law of Mass Action to write equilibrium constant expression for a given reversible reaction.
7. Use Le Chatelier's Principle to predict the direction in which equilibrium will shift if a stress is applied.
8. Calculate equilibrium concentrations/pressures given a chemical equation and initial concentrations/pressures based on equilibrium constant expressions.

Ionic and Acid-Base Equilibria (~15%)

9. Write equations representing acid-base and solubility product equilibria.
10. Calculate concentrations of individual components based on acid base and/or solubility product equilibria and equilibrium constants. .
11. Characterize weak and strong acids/bases, their relevance to pH and to buffer preparation.
12. Relate equilibrium constants and pH to titrations and vice-versa.
13. Use the solubility product calculations to determine when a substance precipitates and apply the understanding of this topic to describe how to separate ions and conduct separations and qualitative analysis based on solubility product equilibrium.

Thermodynamics: (~17.5%)

14. Describe the laws of thermodynamics as they apply to chemical changes.
15. Calculate quantities involving heat, work, internal energy, and enthalpy changes.

16. Calculate changes in entropy(S) and free energy(G) for chemical equations using thermochemistry data.
17. Perform calculations that relate the equilibrium constant to free energy changes in various reversible systems.

Electrochemistry and Oxidation-Reduction: (~15%)

18. Describe electrochemical cells both qualitatively and quantitatively.
19. Perform calculations based on Faraday's Law of Electrolysis.
20. Apply the Nernst equation to non-standard state cells.
21. Complete and balance oxidation-reduction equations in acidic, basic, and non-aqueous media.

Nuclear Chemistry: (~10%)

22. Calculate binding-energy changes for nuclear processes.
23. Use the half-life of a nuclide to calculate quantities based on natural decay.
24. Complete and balance nuclear equations.
25. Describe nuclear chemistry applications including fission, fusion, transmutation of elements, and the use of radioactive tracers.

Descriptive Chemistry: (~15%)

26. Describe the properties and uses of selected materials such as semiconductors, superconductors, ceramics, polymers, and fuels.
27. Write acceptable names and formulae for coordination compounds.
28. Given the name and general shape of a coordination compound draw the structure of the compound.
29. Describe bonding of coordination compounds in terms of valence bond and crystal field theory.
30. Describe the bonding and isomerization in organic compounds.
31. Identify the important functional groups for organic molecules.
32. Name simple organic compounds.
33. Identify and write the structure for specific polymers.

Lab Outline

Lab Component in this Course

Yes

Lab Component Description

Be sure that your outline covers the time spent in class. It can be expressed in weeks, hours, or as a percentage of the course. If in weeks, make sure it covers 16 weeks. You can use week 16 for finals if needed. Put N/A if this is a CCN course.

Lab:

Laboratory:

Safety: (On Going)

1. Describe the hazards of specific classes of chemical substances.

2. Describe the risks involved in selected laboratory operations.
3. Describe and apply the precautions needed to avoid injury in chemistry laboratories.
4. Students will be expected to pass a safety test. Laboratory Content and Relative Experiments: (108 hours)
 - A. Safety Procedures and Resources (~5 %)
 - B. Experiments, which explore Study of Reaction Rates and order of reactions, (~15%)
 - C. Le Chatelier's Principle, Equilibrium Shift and Equilibrium Constant Determination, (~15%)
 - D. K_a/K_b of Acids/Bases, Titration Curves and Buffers, (~15%)
 - E. Qualitative Analysis of Groups 1-5, (~30 %)
 - F. Entropy/Gibb's Free Energy, (~10%)
 - G. Oxidation/Reduction Reactions, (~5%)
 - H. Electrochemical Cells and Electrolysis, (~5%)
 - I. Special Project cooperatively chosen by student and instructor. (As Needed)

DISTANCE AND CORRESPONDENCE EDUCATION CRITERIA AND STANDARDS

CORRECT Please choose all of the delivery methods applicable to this course.

Face to Face Hybrid (requires face to face meetings)

CORRECT Is this course being submitted for Correspondence Education approval?

No

CORRECT Rigor statement: The same standards of course quality shall be applied to distance education as are applied to traditional classroom courses in regard to the course quality judgments made pursuant to the requirements of Section 55002. The same expectations applies to any local course quality determination or review process.

Methods of evaluation and out of class assignments are NOT the same as for a face to face course.

CORRECT If the methods of evaluation differ from a face to face courses, please indicate what the differences are and why they are being used.

If this class were taught as a distance learning course, the labs would have to be done very differently. For that reason, it is not being submitted as a distance learning course. A hybrid course would meet in person for lab and have the lecture content be online.

CORRECT If the face to face course has a lab, field trip, or site visit explain how these components will be performed in the online course. Be sure to identify how the lab component will differ from a homework assignment.

CORRECT How will "regular and substantive interaction" be demonstrated in this course? Title V, section 55204 requires at least TWO of the following. Please choose at least TWO options from the list below.

Assessing or providing feedback on a student's coursework
Providing information or responding to questions about the content of a course or competency

CORRECT If you chose "Other instructional activities approved by the institution's or program's accrediting agency" in the above question, please provide supplemental information.

CORRECT Title V Section 55204 requires regular interaction by instructors.. By checking both boxes below, the instructor is ensuring compliance with Section 55204 of Title V.

Instructor will ensure that students are provided the opportunity for substantive interactions with the student on a predictable and scheduled basis. commensurate with the length of time and the amount of content in the course or comp... Instructor will ensure that the student's academic engagement and success is monitored and ensure that the instructor is responsible for promptly and proactively engaging in substantive interaction with the student when needed on the... or upon request by the student.

CORRECT Purely because of the delivery mode, will you require additional software or hardware beyond basic computer and web browser capabilities?

No

CORRECT If you answered YES to the previous question, please explain below.

CORRECT Federal and state regulations require that all online course materials be made available in an accessible electronic format. By checking both boxes below, the instructor is ensuring compliance with Section 508 of the Rehabilitation Act.

Instructor will ensure the course is 508 compliant using the Course Management System and other tools as needed. Instructor will ensure text and any other courses materials are 508 compliant.

CORRECT A good practice is that section size should be no greater in distance education modes than in regular face to face versions of the course. Will the online section for this course differ from face to face sections?

No

CORRECT If the online section of the course will differ in size from face to face sections, please provide a rationale for the size difference.

—

CURRICULUM TECHNICIAN ONLY

Course Control Number - CB00

eL-Course Control Number

—

CCC000585236

Course Program Status - CB24

1

Repeat Status

Course not repeatable

Repeat Maximum Hours

—

Course Title - CB02

General Chemistry and Chemical Analysis

Course Description

Prerequisites: Successful completion of CHEM B1A or equivalent with a grade of C or better and successful completion of MATHB70 or equivalent with a grade of C or better.

Advisory: Successful completion of MATHB1A or equivalent with a grade of C or better.

Description: This course is a continuation of CHEM B1A. The course is intended to provide a broad foundation of chemical principles for science majors. The major topics considered include the properties of solutions and solids, chemical equilibrium, acids and bases, solubility products, thermodynamics, electrochemistry, and nuclear chemistry. The lab portion of the class will include quantitative measurements based on the same topics.

C-ID: Chemistry (CHEM 120S (when both CHEM B1A and CHEM B1B completed))

Instruction Modes

Status

Laboratory/Studio/Activity

Delayed Interactn - NET

Simultaneous Interact - NET

Lecture/Discussion

Distance Ed - 2 Way Interact

Approved

Board of Trustees Approval Date

Course ID

Feb 20, 2025

3880

Course Version Status

Curriculum Committee Approval Date

Date for Next Review

Approved

Dec 5, 2024

Jul 1, 2030

Time to Next Review

5 years

Course Attributes

Elective (BELE) BC GE Natural Sciences (BC5) BC GE Natural Sciences (GBBN) CSU GE Area B-1 - Phys. Sci. (CBPH)
CSU GE Area B-3 (CBB3) CSU Transferable (CSU) IGETC Area 5 - Lab Course (I5LA) IGETC Area 5 - Physical Sci. (I5PH)
UC Transferable (UC)

Effective Start Date

May 19, 2025

Division

Physical Science

College Name

[Bakersfield \(BC\)](#)

Academic Group/College

[Bakersfield College](#)

Hidden Fields

End Term

—

Approval Code

[Credit - Degree Applicable](#)

eL-Department

[Chemistry](#)

eL-Course Crosswalk CRS-DEPT-NAME (CB19)

[CHEM](#)

National Taxonomy Code

[40.0501](#)

Allow Students to Audit Course

[Yes](#)

eL- Short Course Title

—

CAN Code - CB14 CAN Seq Code - CB15

—

Course Crowsswalk CRS-DEPT-NAME - CB19

[CHEM](#)

Mode of Delivery

—

Credit Type

[Institutional](#)

Curriculum Id

[CHEMB1B](#)

eL-Course Code

[CHEMB1B](#)

eL-Course Subject

[CHEM](#)

eL-Course Number

[B1B](#)

Course Level

—

Course Code Suffix

—

Course Family

—

eL-Academic Career —

—

eL-Campus

—

Academic Term Type

—

eL-Mode of Delivery

—

eL-Academic Group

—

Prerequisite Waiver Code

—

COURSE-DEPARTMENT-NUMBER - CB01

[839d4982-f7cd-46aa-ae37-ac4e23817430](#)

eL-Course Crosswalk CRS-NUMBER (CB20)

[B1B](#)

National Taxonomy Name

[Chemistry, General](#)

Course Crowsswalk CRS-NUMBER - CB20

[B1B](#)

External Course ID Stacked Courses Only — Stacked Courses Only Rationale —

No

No

—

Estimated Enrollment

eLearning Platform —

—

—

Estimated Section Size — Default Section Size — Maximum Section Size — Minimum Section Size —

—

—

—

—

Total Percent Weight (%)

—

eL-Distance Education Approved

eL-Distance Education Approved

Yes

Yes

Program Name

Biology Associate in Science for Transfer, Chemistry Associate in Science for Transfer, Engineering Associate of Science, Geology Associate in Science for Transfer, Human Biology Associate of Science, Nutrition and Dietetics Associate in Science for Transfer, Nutrition and Dietetics Associate in Science for Transfer (In Development), Research Laboratory Technology Associate of Science, Research Laboratory Technology Bachelor of Science, Research Laboratory Technology Bachelor of Science (In Development), UCTP Chemistry, UCTP Physics, UCTP Physics (In Development)

Assessment Related Requirement

—

Will this course be part of a program?

Yes

eL-Status

Active